

These notes can be obtained at: <http://www.ndsu.nodak.edu/instruct/grcook/chem342/notes.shtml>

● Chapter 12: Mass Spectrometry and Infrared Spectroscopy

How do we determine the structure of organic molecules?

Probe the physical properties

Elemental Analysis (combustion analysis)

Atomic composition (relative ratios)

Empirical formula

Mass Spectrometry

Molecular formula

Elemental identification (isotopes)

Structural features (from fragments)

Infrared (Vibrational) Spectroscopy

Functional Group Identities

Ultraviolet (electronic) Spectroscopy

Conjugated pi-systems

Nuclear Magnetic Resonance Spectroscopy

Atom connectivity

Functional groups

stereochemistry

X-Ray Crystallography

3D position of atoms

Elemental Analysis

Provides empirical formula

Percent Composition
by Mass

C - 63.31%
H - 6.28%
Cl - 16.99%
N - 13.42%

Assume 100g sample

⇒ %composition = grams

Moles of C = $\frac{63.31 \text{ g}}{12.011 \text{ g/mol}} = 5.27$	$5.27 / 0.48 = 11$
Moles of H = $\frac{6.28 \text{ g}}{1.008 \text{ g/mol}} = 6.23$	$6.23 / 0.48 = 13$
Moles of N = $\frac{13.42 \text{ g}}{14.007 \text{ g/mol}} = 0.96$	$0.96 / 0.48 = 2$
Moles of Cl = $\frac{16.99 \text{ g}}{35.45 \text{ g/mol}} = 0.48$	$0.48 / 0.48 = 1$

⇒ Empirical Formula = $\text{C}_{11}\text{H}_{13}\text{N}_2\text{Cl}$

Units of Unsaturation

The number of pi-bonds and rings in the molecule

Saturated hydrocarbon has the general formula $\text{C}_n\text{H}_{2n+2}$

Every pi-bond or ring removes 2 H's – each loss of 2 H's is an unsaturation

Difference in the number of H's between the saturated and the unsaturated molecule

Every halogen replaces one H and must be counted

Every nitrogen adds one H and must be counted

General Formula for Units of Unsaturation

$$\text{UN} = \frac{(2n+2) - \#H - \#X + \#N}{2}$$

$2n+2$ is the number of H's if completely saturated
#H is the actual number of H's in the molecule
#X is the actual number of halogens
#N is the actual number of nitrogens
divide by 2 because each unit of unsaturation removes 2 H's from the molecule

$$\text{Empirical Formula} = \text{C}_{11}\text{H}_{13}\text{N}_2\text{Cl} \quad \text{UN} = \frac{(2 \cdot 11 + 2) - 13 - 1 + 2}{2} = 6$$